Exemples de la composition de quelques sels doubles.

Noms.	Formules.
Carbonas magnesico-calcicus, bitterspat, dolomic	$\ddot{C}a\,\ddot{C}^2 + \ddot{M}g\,\ddot{C}^2$
Fluosilicias ammonicus	$(3\dot{N}H^6 + 2\ddot{S}i) + (3\dot{N}H^6 + 3\ddot{F})$
hydricus	$3\ddot{F}Aq^2 + 2\ddot{S}i^2\ddot{F}^3$
kalicus	$3\ddot{K}\ddot{F} + 2\ddot{S}i^2\ddot{F}^3$
Oxalas ammonico-cupricus	$2\dot{N}H^6\bar{O}Aq^2 + \ddot{C}u\bar{O}^2Aq^2$
triammonico-cupricus. . . .	$2(3\dot{N}H^6 + \bar{O}Aq^3) + \ddot{C}u^3\bar{O}^2Aq^6$
Oxalas kalico-cupricus c. aq. var. 1:ma.	$\ddot{K}\,\bar{O}^2Aq + \ddot{C}u\,\bar{O}^2Aq$
var. 2 : da.	$\ddot{K}\,\bar{O}^2Aq^2 + \ddot{C}u\,\bar{O}^2Aq^2$
natrico-cupricus	$\ddot{N}a\,\bar{O}^2Aq + \ddot{C}u\,\bar{O}^2Aq$
Murias ammonico-ferrosus	$2\dot{N}H^6\,\ddot{M} + \ddot{F}e\,\ddot{M}^2$
hydrargyricus . . .	$2\dot{N}H^6\,\ddot{M}Aq + \ddot{H}g\,\ddot{M}^2$
platinicus.	$2\dot{N}H^6\ddot{M}Aq + \ddot{P}t\,\ddot{M}^2$
kalico-platinicus	$\ddot{K}\,\ddot{M}^2 + \ddot{P}t\,\ddot{M}^2$
natrico-platinicus	$\ddot{N}\,\ddot{M}^2 + \ddot{P}t\,\ddot{M}^2$
Hydro-carbonas cupricus.	$\ddot{C}u\,Aq^2 + 2\ddot{C}u\,\ddot{C}^2$
magnesicus. . . . . (magnesia alba).	$\ddot{M}g\,Aq^8 + 3\ddot{M}g\,\ddot{C}^2$
zincicus.	$\ddot{Z}n\,Aq^6 + 3\ddot{Z}n\,\ddot{C}$
Murio-carbonas plumbicus.	$\ddot{P}b\,\ddot{M}^2 + \ddot{P}b\,\ddot{C}^2$
Sulphas aluminico-ammonicus	$\dot{N}H^6\ddot{S} + \ddot{A}l\,\ddot{S}^3$
kalicus	$\ddot{K}\,\ddot{S}^2 + 2\ddot{A}l\,\ddot{S}^3$
cum aquâ . .	$\ddot{K}\,\ddot{S}^2 + 2\ddot{A}l\,\ddot{S}^3 + 48Aq.$
natricus	$\ddot{N}a\,\ddot{S}^2 + 2\ddot{A}l\,\ddot{S}^3$
ammonico-cupricus.	$2\dot{N}H^6\ddot{S}Aq^2 + \ddot{C}u\,\ddot{S}^2Aq^{10}$
triammonico-cupricus. . . . (cuprum ammoniacum)	$4(3\dot{N}H^6 + \ddot{S}) + \ddot{C}u^3\ddot{S}^2Aq^6$
ammonico-kalicus	$\ddot{K}\,\ddot{S}^2 + 2\dot{N}H^6\ddot{S}Aq^2$